

Review C2.7 Electrolysis

<i>Can you...?</i>	😊	😐	😞
Describe what electrolysis is and what it does.			
State the type of compound that can be used as an electrolyte			
Explain why the electrolyte must be molten or in solution for electrolysis to work			
Describe which ions move to which electrode.			
Explain what then happens to ions at that electrode, in terms of electrons.			
Describe how electrolysis is used to electroplate objects with copper or silver plating.			
Remember OIL RIG and describe what it means in terms of electrons			
Recall that if there's a mix of ions, the products formed depend on the reactivity of elements involved.			
[HT] Represent reactions at electrodes using half equations. For example: $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$ or $2\text{Cl}^- - 2\text{e}^- \rightarrow \text{Cl}_2$			
Describe how aluminium is manufactured by electrolysis			
Explain why cryolite is needed in the electrolysis of aluminium oxide.			
Explain why the carbon electrodes in the electrolysis of aluminium must be replaced often.			
Describe the details of the electrolysis of sodium chloride solution (brine)			
Explain why the products of brine electrolysis are useful reagents in the chemical industry – particularly soap, bleach and plastics.			