

Review C2.4 Rates of reaction

| Can you...? | 😊 | 😐 | 😞 |
|---|---|---|---|
| Look at a graph and work out the rate of reaction from products forming. | | | |
| Describe the changing rate of a reaction by looking at a graph | | | |
| Evaluate the advantages and disadvantages of using catalysts in industry | | | |
| Can you calculate the rate of a reaction using this: Rate of reaction = $\frac{\text{Amount of reactant used/ product formed}}{\text{Time}}$ | | | |
| Name 5 factors that would affect the rate of a reaction | | | |
| Describe collision theory in terms of particles and energy | | | |
| Recall the name of the energy needed to be overcome to start a reaction | | | |
| Explain how each factor would affect the rate of reaction using collision theory | | | |
| State what a catalyst is and what it does | | | |
| [HT] Recall the unit of concentrations of solutions | | | |
| [HT] Recall that equal volumes of gases at the same temperature and pressure have the same number of molecules | | | |