

Review C2.5 Exothermic/Endothermic Reactions

<i>Can you...?</i>	😊	😐	☹️
State that when chemical reactions occur, energy is transferred to or from the surroundings.			
State what an exothermic reaction is in terms of energy and give examples.			
Define an endothermic reaction in terms of energy and give examples.			
Recall that if a reversible reaction is exothermic in one direction, it is endothermic in the opposite direction.			

Review C2.6 Acids, bases and salts

<i>Can you...?</i>	😊	😐	☹️
Use the state symbols in equations - (s), (l), (g) and (aq).			
Describe how soluble salts can be made by reacting acids with metals, insoluble bases and alkalis			
Describe how salt solutions can be crystallised to produce solid salts.			
Insoluble salts can be made by mixing certain salts in solution (precipitate formed)			
Describe how precipitation can be used to remove unwanted ions from solutions, e.g. in treating water for drinking.			
Describe the difference between a base (metal oxides) and an alkali (metal hydroxides)			
Name the salts that HCl, HNO ₃ , H ₂ SO ₄ produce			
State that ammonia dissolves in water to produce an alkaline solution. It is used to produce ammonium salts, which are important as fertilisers.			
Recall that the pH scale is a measure of the acidity or alkalinity of a solution.			
Describe an acid as releasing H ⁺ ions in solution.			
Describe an alkali as releasing OH ⁻ ions in solution.			
In neutralisation reactions, hydrogen ions react with hydroxide ions to produce water. Represent this reaction with the equation: $\text{H}^+_{(\text{aq})} + \text{OH}^-_{(\text{aq})} \rightarrow \text{H}_2\text{O}_{(\text{l})}$			